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Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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Chapter Exam Instructions. Choose your answers to the questions and click 'Next' to see the next set of questions.

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2- a) 2 mol HgO 2 mol Hg & 2 mol HgO 1 mol O<sub>2</sub>...

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