

## Chemquest 12 Quantum Numbers Answers

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Explain giving reasons, which of the following sets of quantum numbers are not possible? [Quantum Numbers \(tagalog\) | NOW I KNOW](#)

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l - Secondary Quantum Number/Orbital Shape Quantum number: represents the shape of the orbital- s, p, f, d. l is a range of n-1. ml - Magnetic quantum number: represents the number of orbits possible. M l is a range of l. ms - Spin Quantum number: represents the electron and its spin. Two possibilities +1/2,

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When l = 0 (s cloud), 1 (p cloud), 2 (d cloud), 3 (f cloud) Third - Magnetic Quantum number (ml) = location or spatial orientation ml = - l to l Fourth - Spin Quantum number (ms) = electron spin ms = +1/2 or -1/2. 2. Name the orbitals described by the following quantum numbers a. n = 3, l = 0 3s. c. n = 3, l = 2 . 3d

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