

Chemistry 1a Net Ionic Answer Key

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ionic: $H^+(aq) + Cl^-(aq) + K^+(aq) + OH^-(aq) \rightarrow K^+(aq) + Cl^-(aq) + H_2O(l)$ net ionic: $H^+(aq) + OH^-(aq) \rightarrow H_2O(l)$ 3. $H_2SO_4(aq) + Mg(OH)_2(aq) \rightarrow MgSO_4(aq) + 2H_2O(l)$ ionic: $2H^+(aq) + SO_4^{2-}(aq) + Mg^{2+}(aq) + 2OH^-(aq) \rightarrow Mg^{2+}(aq) + SO_4^{2-}(aq) + 2H_2O(l)$ net ionic: $H^+(aq) + OH^-(aq) \rightarrow H_2O(l)$ 4. $NH_4OH(aq) + HCl(aq) \rightarrow NH_4Cl(aq) + H_2O(l)$ ionic: $NH_4^+(aq) + OH^-(aq) + H^+(aq) + Cl^-(aq) \rightarrow NH_4^+(aq) + Cl^-(aq) + H_2O(l)$ net ionic: $H^+(aq) + OH^-(aq) \rightarrow H_2O(l)$

9-1 GCSE Chemistry Ionic Equations Answers

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Chemistry 1a Net Ionic Answer CHEM 1A Review: Experimental Errors, Significant Figures, and Net Ionic Equations Lab Section Name Date 5. Iron III nitrate and potassium sulfide 6. strontium acetate and phosphoric acid - 7. aluminum sulfate and ammonium sulfide 8. barium hydroxide and zinc(II)

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$2K^+(aq) + 2Cl^-(aq) + Pb^{2+}(aq) + 2NO_3^-(aq) \rightarrow 2K^+(aq) + 2NO_3^-(aq) + PbCl_2(s)$ Net Ionic Equation: $2Cl^-(aq) + Pb^{2+}(aq) \rightarrow PbCl_2(s)$ Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in aqueous solution.

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Chemistry 1a Net Ionic Answer Answer to 1a. Identify whether each species functions as a Brønsted-Lowry acid or a Brønsted-Lowry base in this net ionic equati... Solved: 1a. Identify Whether Each Species Functions As A B ... $2K^+(aq) + 2Cl^-(aq) + Pb^{2+}(aq) + 2NO_3^-(aq) \rightarrow 2K^+(aq) + 2NO_3^-(aq) + PbCl_2(s)$ Net Ionic

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